

## suite EX 6

$$\underline{m_{\text{sac}} = 960 \text{ g.}}$$

b) 100 g de solution Saccharose.  $\longrightarrow$  80 g de saccharose pur.

$$960 \text{ g} \longrightarrow m'_{\text{sac}} = \frac{80 \times 960}{100} = 768 \text{ g.}$$

$$\underline{m'_{\text{sac}} = 768 \text{ g.}}$$

c) La molarité de la solution :  $C$ .

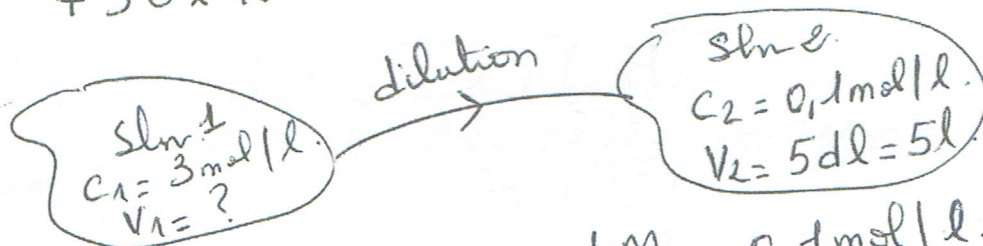
$$C = \frac{n_{\text{sac. pur.}}}{V_{\text{solution}}} \quad (\text{mol./l.})$$

$$n_{\text{sac}} = \frac{m'_{\text{sac}}}{M_{\text{sac}}} = \frac{768}{(12 \times 12) + (1 \times 22) + (11 \times 16)} = \frac{768}{342}$$

$$\underline{n_{\text{sac}} = 2,25 \text{ mol.}}$$

$$C = \frac{2,25}{750 \times 10^{-3}} = \underline{3 \text{ mol./l} = C.}$$

d)



Solution décimolaire  $\Rightarrow C = 0,1 \text{ M} = 0,1 \text{ mol/l.}$

Loi de dilution :  $C_1 V_1 = C_2 V_2 \Rightarrow V_1 = \frac{C_2 V_2}{C_1}$

$$V_1 = \frac{0,1 \times 5}{3} = 0,167 \text{ l} = 167 \text{ ml}$$

$$\underline{V_1 = 167 \text{ ml.}}$$